

Mono-Atomic Ions & Ionic Compound Formation

ANSWERS

1	Rb^{+1}	Cl^{-1}	RbCl	rubidium chloride
2	Cs^{+1}	S^{-2}	Cs_2S	cesium sulfide
3	Fr^{+1}	P^{-3}	Fr_3P	francium phosphide
4	Be^{+2}	Br^{-1}	BeBr_2	beryllium bromide
5	Mg^{+2}	O^{-2}	MgO	magnesium oxide
6	Ca^{+2}	N^{-3}	Ca_3N_2	calcium nitride
7	Sr^{+2}	I^{-1}	SrI_2	strontium iodide
8	Ba^{+2}	S^{-2}	BaS	barium sulfide
9	Al^{+3}	P^{-3}	AlP	aluminum phosphide

D	Li	Br	LiBr	lithium bromide
10	Na ⁺¹	O ⁻²	Na ₂ O	sodium oxide
11	K ⁺¹	P ⁻³	K ₃ P	potassium phosphide
12	Rb ⁺¹	I ⁻¹	RbI	rubidium iodide
☺	Cs ⁺¹	O ⁻²	Cs ₂ O	cesium oxide
14	Fr ⁺¹	P ⁻³	Fr ₃ P	francium phosphide
15	Be ⁺²	Cl ⁻¹	BeCl ₂	beryllium chloride
16	Mg ⁺²	S ⁻²	MgS	magnesium sulfide
17	Ca ⁺²	P ⁻³	Ca ₃ P ₂	calcium phosphide
E	Sr	S	SrS	strontium sulfide
18	Ba ⁺²	P ⁻³	Ba ₃ P ₂	barium phosphide
19	Al ⁺³	F ⁻¹	AlF ₃	aluminum fluoride
20	Li ⁺¹	O ⁻²	Li ₂ O	lithium oxide
21	Na ⁺¹	N ⁻³	Na ₃ N	sodium nitride
22	K ⁺¹	Cl ⁻¹	KCl	potassium chloride
23	Rb ⁺¹	S ⁻²	Rb ₂ S	rubidium sulfide
F	Cs	P	Cs ₃ P	cesium phosphide
24	Fr ⁺¹	Br ⁻¹	FrBr	francium bromide
25	Be ⁺²	O ⁻²	BeO	beryllium oxide