SF Practice

How many SF are in each measurement? Name:

14.56 mL	0.450 g	1.0000008 cm	0.000008 cm	6.36×10^4 kg	300 ounces	100 m
4507 joules	0.45 g	6.5×10^{23} atoms	2.4000 m	2.4001 m	4500.1 grams	100. m
0.4570 joules	15 meters	5.5556×10^2 kPa	0.1 amt	0.11 atm	0.110 atm	100.0 m
0.15 meters	1.5 meters	5.000×10^9 g	7°C	14°C	14.000°C	14.000004°C
0.150 meters	1.0 cm	0.1 cm	0.10 cm	273 Kelvin	0.008 grams	7.00×10^3 °C

Questions. Answers into the six boxes, work on white paper, staple onto this handout.

1. Calculate density of an unknown metal with mass of 76.12462 g, and volume of 14.300 mL.

2. Using table S, what metal is that from question 1? Write the name, symbol and atomic number.

3. Write the density formula. If the density of water is 1.000 g/mL, what is the volume of 239 g H_2O?

4. Which has a greater density, 20.0 grams of aluminum or 14.0 grams of aluminum? Explain.

5. Write the symbols for both water and gold. One is an element, one is a compound, which is which?

6. Write the density formula again. What is the volume of 239 grams of gold?

PUT ANSWERS INTO THESE BOXES. Staple your work to this handout. No work + no credit.

1	2	3
4	5	6