Reactions HW #2 - Single Replacement Reactions Name: Key

Write balanced chemical reactions for each of these, with phase symbols. If there is no reaction, Write the symbols for the reactants, an arrow and an "X". A solution means AQUEOUS.

Get Table J

1	Zinc reacts with hydrochloric acid solution	$Zn_{(s)} + 2HCl_{Aq} \longrightarrow ZnCl_{2\,Aq} + H_{2(g)}$
2	Iron (II) nitrate solution plus silver metal	$Fe(NO_3)_{2\,Aq} + Ag \longrightarrow X$ no reaction
3	Fluorine gas mixed with sodium bromide solution	$F_{2(g)} + 2NaBr_{Aq} \longrightarrow 2NaF_{Aq} + Br_{2(l)}$
4	Gold (III) chloride solution with magnesium metal	$2AuCl_{3\,Aq} + 3Mg_{(s)} \longrightarrow 3MgCl_{2\,Aq} + 2Au$
5	Copper (II) sulfate solution with silver metal	$CuSO_{4\,Aq} + Ag_{(s)} \longrightarrow X$ no Reaction
6	Bromine liquid into ammonium iodide solution	$Br_{2\,L} + 2NH_4I_{Aq} \rightarrow 2NH_4Br_{Aq} + I_2$
7	Ammonium fluoride solution with chlorine gas	$NH_4F_{Aq} + Cl_{2(g)} \longrightarrow X$ no reaction
8	Lithium hydroxide solution with titanium metal	$LiOH_{Aq} + Ti_{(s)} \longrightarrow X$ no reaction
9	Barium hydrogen carbonate solution and lithium metal	$Ba(HCO_3)_{2\,Aq} + 2Li_{(s)} \rightarrow 2LiHCO_{3\,Aq} + Ba_{(s)}$
10	Potassium sulfate solution with lead metal	$K_2SO_{4\,Aq} + Pb_{(s)} \longrightarrow X$ no reaction
11	Aluminum metal into nickel (II) chlorate solution	$2Al_{(s)} + 3Ni(ClO_3)_{2\,Aq} \rightarrow 2Al(ClO_3)_{3\,Aq} + 3Ni_{(s)}$