Reactions HW #3 - Double Replacement Reactions Name: key

Write balanced equations with PHASES, which are mandatory. Use the Solubility Guidelines Table (table F). One of these is not a reaction, will you find it? If it is a "no reaction", you will still balance the equation and indicate both products as AQ. Do not just write NO REACTION, that is not enough.

Sodium carbonate + zinc chlorate solutions react

$Na_2CO_{3\,Aq} + Zn(ClO_3)_{2\,Aq} \rightarrow ZnCO_{3(s)} + 2NaClO_{3\,Aq}$

Copper (II) sulfate + calcium chloride solutions react

$CuSO_{4\,Aq} + CaCl_{2\,Aq} \rightarrow CuCl_{2\,Aq} + CaSO_{4(s)}$

Potassium hydroxide + lead (IV) nitrate solutions react

$4KOH_{Aq} + Pb(NO_3)_{4\,Aq} \rightarrow Pb(OH)_{4(s)} + 4KNO_{3\,Aq}$

Silver hydrogen carbonate + Iron (II) bromide solutions react

$2AgHCO_{3\,Aq} + FeBr_{2\,Aq} \rightarrow 2AgBr_{(s)} + Fe(HCO_3)_{2\,Aq}$

Barium hydroxide + lithium sulfate solutions react

$Ba(OH)_{2\,Aq} + Li_2SO_{4\,Aq} \rightarrow BaSO_{4(s)} + 2LiOH_{Aq}$

Ammonium phosphate + tin (II) acetate solutions react

$2(NH_4)_3PO_{4\,Aq} + 3Sn(C_2H_3O_2)_{2\,Aq} \rightarrow 6NH_4C_2H_3O_{2\,Aq} + Sn_3(PO_4)_{2(s)}$

Calcium chromate + sodium sulfide solutions react

$CaCrO_{4\,Aq} + Na_2S_{Aq} \rightarrow CaS_{(s)} + Na_2CrO_{4\,Aq}$

Strontium acetate + lithium carbonate solutions react

$Sr(C_2H_3O_2)_{2\,Aq} + Li_2CO_{3\,Aq} \rightarrow SrCO_{3(s)} + 2LiC_2H_3O_{2\,Aq}$

Ammonium phosphate + rubidium nitrate solutions react

$(NH_4)_3PO_{4\,Aq} + 3RbNO_{3\,Aq} \rightarrow$ "no Reaction"

$3NH_4NO_{3\,Aq} + Rb_3PO_{4\,Aq}$ 2 AQ products