

How to use this handout for your review.

Six tables, front three, back three.

One set of cards, in three sets, color coded in three envelopes marked AB and C.
One set of answer cards in separate envelopes travels to the table with the question cards.

Teacher assigns four students to table, mix appropriately.

Two teams work on all eight problems, but work only with your partner until answer check, then all 4 students can participate together.

Once all eight questions get checked and solved, put all questions back into proper envelopes.

Teacher moves envelope packets table to table.

Students choose a new partner from the four people at the table. Solve all eight new problems the same way with your new partner. All four participate in answer checking.

Wrap up cards again, teacher again distributes cards to tables, students choose the last possible pairing, repeat.

Each student should manage 18 perfect answers by the end. Teacher is last chance person to solve any questions.

Wrap up with discussion about how groups went, how this review compares to blackboard problems, etc.

Wrap up all cards in proper envelopes and return to teacher.

Teacher must have seating chart for overhead before class to save time.

Students (in chem) need paper, pen, and reference tables (maybe calculators, depending upon the problem sets).

Print question cards on 2 different color card stock, all answers can be white, or color coded.

Make sure every student gets an answer card at each table, 4 per table needed for time saving.

A

Name this compound



A

Name this compound



A

Name this compound



A

What is the formula for

Nitrogen trichloride

A

What is the formula for

Dinitrogen pentoxide

A

What is the formula for

Sulfur trioxide

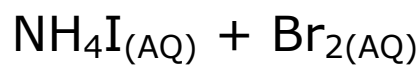
A

Balance this reaction



A

Balance this reaction



B

What is the formula for
Vanadium (V) sulfide

B

Name this compound
 PI_3

B

What is the formula for
Zirconium phosphate

B

Name this compound
 C_3H_8

B

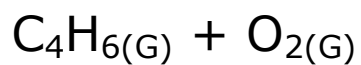
What is the formula for
Palladium (IV) hydroxide

B

Name this compound
 $\text{Hg}_2(\text{SCN})_2$

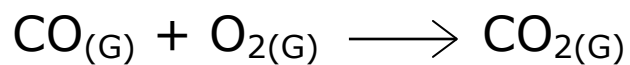
B

Balance this reaction



B

Balance this reaction



C

What is the formula for



C

Name this compound



C

What is the formula for

Tin (II) bromide

C

What is the formula for

Niobium (V) oxalate

C

Name this compound



C

Write the formula
from this name

Selenium tetrafluoride

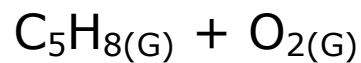
C

Balance this reaction



C

Balance this reaction

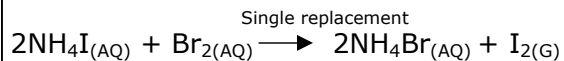
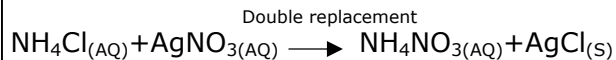


Directions for PRE–Stoichiometry Review

1. Sit where assigned on overhead. Only bring paper, pen, and reference tables with you, put all other materials in the back of the room
2. Choose one partner at your table for the first round of questions. Realize that there are three rounds, you will partner up with each of the students at your table during this class period.
3. Open the questions envelope, share the cards with the other team. Do ALL the questions with your one partner. When every one is done with the set of questions, each of you takes an answer card and each of you checks your work. Grade yourself and count how many correct you get.
4. When instructed, put cards back into the envelopes, trade the envelopes, in order, with the next table.
5. Get a new partner from your table, do all questions, check, grade yourself again.
6. Repeat with last set of questions again.
7. End with class discussion, was this helpful for review, was this more “fun” than practicing individually, etc.
8. This will not be graded, the grade is for you and your mind only.

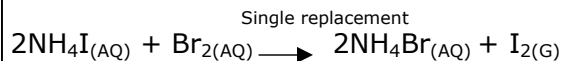
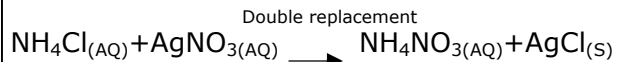
Answers Set A

Manganese (VII) cyanide NCl_3
 Titanium (IV) hydroxide N_2O_5
 Cobalt (III) fluoride SO_3



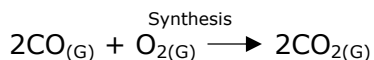
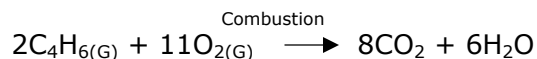
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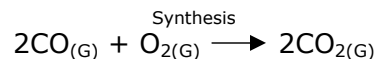
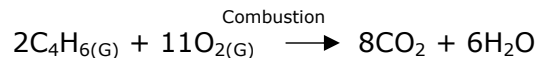
Answers Set B

Phosphorous tri-iodide V_2S_5
 Tricarbon octahydride (propane) $\text{Zr}_3(\text{PO}_4)_4$
 Mercury (II) thiocyanate $\text{Pd}(\text{OH})_4$



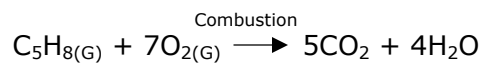
Answers Set B

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 Tricarbon octahydride (propane) $\text{Zr}_3(\text{PO}_4)_4$
 Mercury (II) thiocyanate $\text{Pd}(\text{OH})_4$



Answers C

Bismuth (III) nitride SeF_4
 Nitrogen monoxide $\text{Nb}_2(\text{C}_2\text{O}_4)_5$
 Cobalt (II) bromide SnBr_2



Answers C

Bismuth (III) nitride SeF_4
 Nitrogen monoxide $\text{Nb}_2(\text{C}_2\text{O}_4)_5$
 Cobalt (II) bromide SnBr_2

