

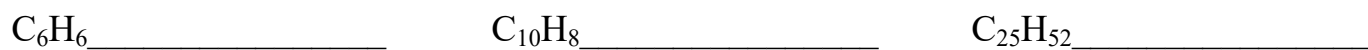
Directions: These assignments are designed to challenge your memories with problems from the “past” that are not allowed to be forgotten. You can use your notes or the BASICS, but please don’t just copy. These will be graded, but if you copy yourself to a 10/10, will you really know it, will you really be challenging yourself, will you really be learning? What’s the point? Learning is fundamental to being in high school. Above all else, don’t be evil.

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To answer each of these three questions, you must calculate the molar mass (on the left side) and then the percent composition by mass (on the right) for each compound, then multiply the given mass by the proper decimals. Watch SF and units.

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1. If you have 36.25 grams of magnesium sulfate, how many grams are just magnesium?
  
  
  
  
  
  
  
  
  
  
  2. If you have 618 grams of sodium hypochlorite (this is the stuff called “chlorine” that people put into pools), how many grams are chlorine? How many grams are sodium?
  
  
  
  
  
  
  
  
  
  
  3. If you have 1530 grams of calcium permanganate, how many grams are calcium? How may grams are manganese? How many grams are just oxygen?

4. Write out the EMPRICIAL FORMULAS for these six compounds



5. Balance these equations (write in a "1" or leave those dashes blank, either is fine)  
~ ADD Phase symbols everywhere you can

