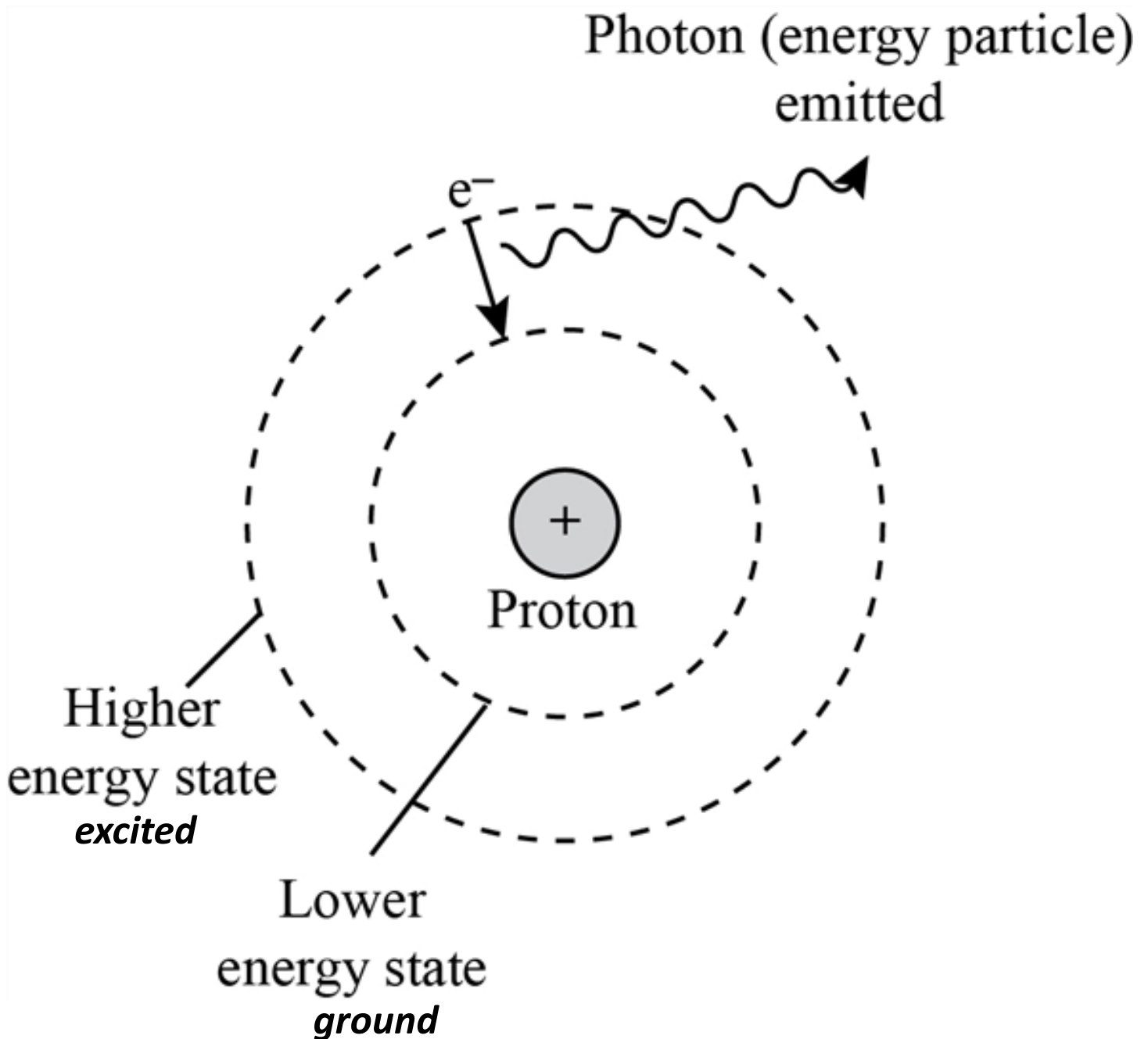
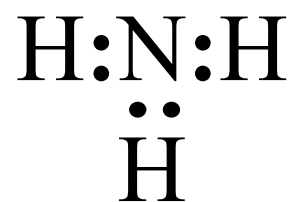
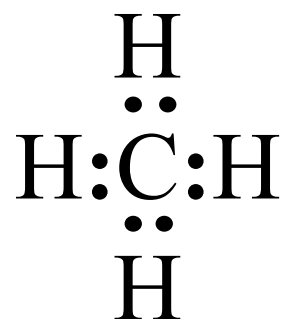
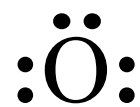
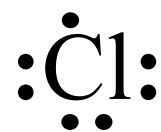
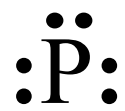


# Spectra Emission



When an electron absorbs energy it will “jump” and move to a higher than normal energy level (excited state). An electron will go to a “higher energy” shell. When that electron returns to the normal, or ground state, it will emit the energy it gained.

**That energy emitted is the spectra.**



Mass of candle to start \_\_\_\_\_ g

Burn candle for 15 minutes

Candle mass at the end \_\_\_\_\_ g

Aluminum foil, rip up into small pieces  
— not into balls.

They have to fit into the test tube.

Pour acid first, add aluminum,  
then “shake” as shown.

$\text{HCl}_{(\text{AQ})}$  is strong enough to hurt you,  
be careful.