

Chemical Reaction Round Robin Review

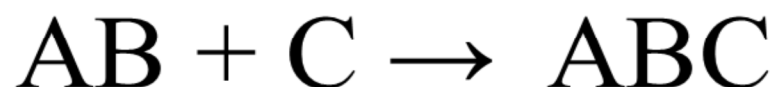
Set up the 12 cards in the back, answers taped onto the back of each card.

Each pair of students will have 3 minutes to answer the two questions,
then move on.

This exercise is a power point as well, on the Reactions Page
of Arbuiso.com

Print on card stock, but ANSWERS on paper.

1. What type of reaction is shown?

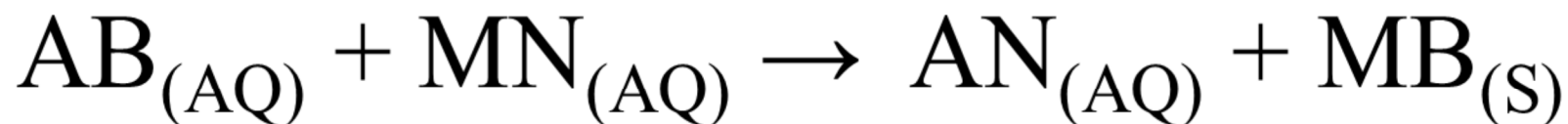


2. What type of reaction is shown?

Balance this too.

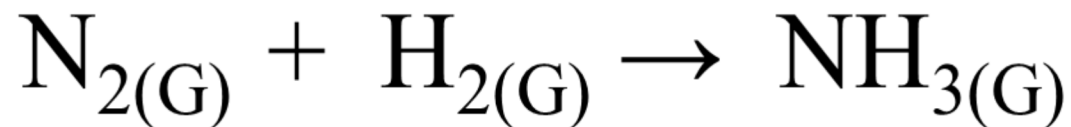


3. What type of reaction is shown?



4. What type of reaction is shown?

Balance this too.



5. What type of reaction is shown?

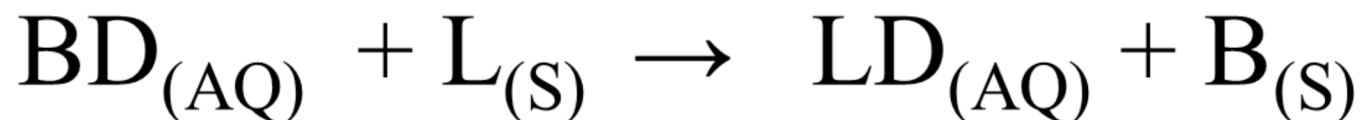


6. What type of reaction is shown?

Balance this too.



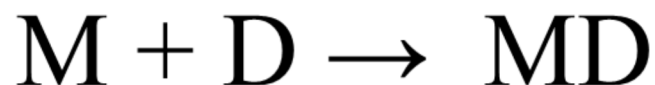
7. What type of reaction is shown?



8. What type of reaction is shown?
Balance this too.

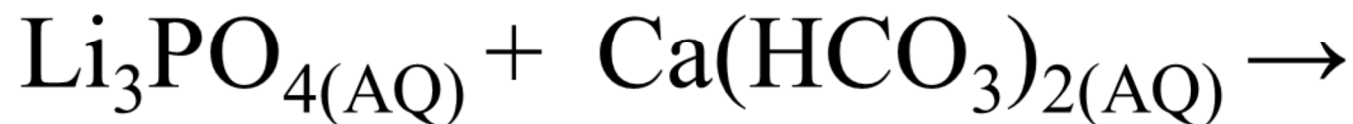


9. What type of reaction is shown?



10. What type of reaction is shown?

Balance this too.



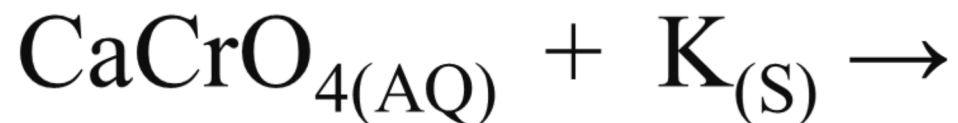
11. What type of reaction is shown?

octane and oxygen forms

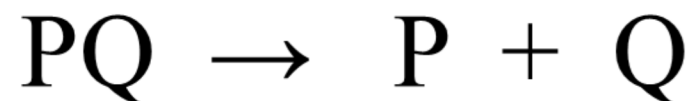
water and carbon dioxide

12. What type of reaction is shown?

Balance this too.



14. What type of reaction is shown?

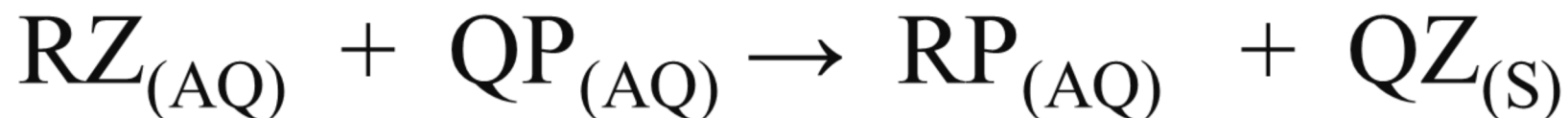


15. What type of reaction is shown?

Balance this too.



16. What type of reaction is shown?

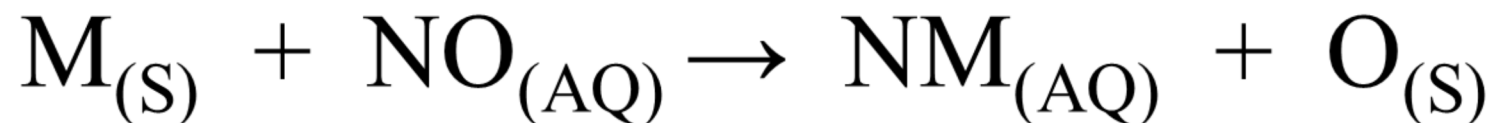


17. What type of reaction is shown?

Balance this too.



18. What type of reaction is shown?

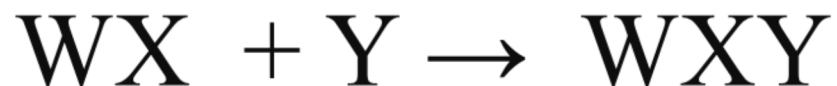


19. What type of reaction is shown?

Balance this too.

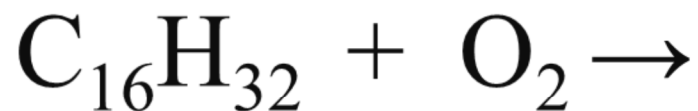


20. What type of reaction is shown?



21. What type of reaction is shown?

Balance this too.

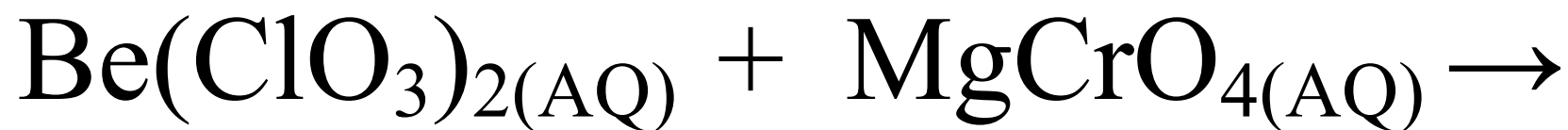


22. What type of reaction is shown/

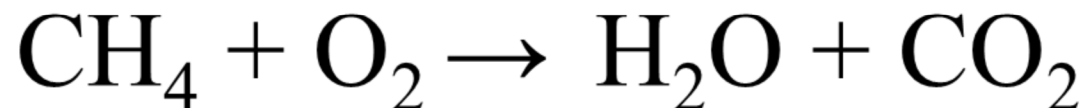


23. What type o reaction is shown?

Balance this too.



24. What type of reaction is shown?



25. What type of reaction is shown?

Balance this too.



1	$\text{AB} + \text{C} \rightarrow \text{ABC}$	synthesis
2	$2\text{H}_2\text{O}_2 \rightarrow \text{O}_2 + 2\text{H}_2\text{O}$	(decomp)
3	$\text{AB}_{(\text{AQ})} + \text{MN}_{(\text{AQ})} \rightarrow \text{AN}_{(\text{AQ})} + \text{MB}_{(\text{S})}$	double replacement
4	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$	synthesis

5	$XYZ \rightarrow XY + Z$ decomposition
6	$2C_6H_{10} + 17O_2 \rightarrow 12CO_2 + 10H_2O$ combustion

7	$BD_{(AQ)} + L_{(S)} \rightarrow LD_{(AQ)} + B_{(S)}$ single replacement
8	$3NH_4OH_{(AQ)} + Al(NO_3)_3_{(AQ)} \rightarrow 3NH_4NO_3_{(AQ)} + Al(OH)_3_{(S)}$ double replacement

9	$M + D \rightarrow MD$ <p style="text-align: right;">synthesis</p>
10	$2\text{Li}_3\text{PO}_{4(\text{AQ})} + 3\text{Ca}(\text{HCO}_3)_{2(\text{AQ})} \rightarrow \text{Ca}_3(\text{PO}_4)_{2(\text{S})} + 6\text{LiHCO}_{3(\text{AQ})}$ <p style="text-align: center;">double replacement</p>
11	<p style="text-align: center;">Octane & oxygen form water & carbon dioxide</p> <p style="text-align: center;">combustion</p>
12	$\text{CaCrO}_{4(\text{AQ})} + 2\text{K}_{(\text{S})} \rightarrow \text{K}_2\text{CrO}_{4(\text{AQ})} + \text{Ca}_{(\text{S})}$ <p style="text-align: right;">single replacement</p>

14	$\text{PQ} \rightarrow \text{P} + \text{Q}$ <p>decomposition</p>
15	$3\text{Sr} + \text{N}_2 \rightarrow \text{Sr}_3\text{N}_2$ <p>synthesis</p>
16	$\text{RZ}_{(\text{AQ})} + \text{QP}_{(\text{AQ})} \rightarrow \text{RP}_{(\text{AQ})} + \text{QZ}_{(\text{S})}$ <p>double replacement</p>
17	$\text{C}_5\text{H}_{12} + 8\text{O}_2 \rightarrow 5\text{CO}_2 + 6\text{H}_2\text{O}$ <p>combustion</p>

18	$\text{M}_{(\text{s})} + \text{NO}_{(\text{aq})} \rightarrow \text{NM}_{(\text{aq})} + \text{O}_{(\text{s})}$ <p style="text-align: center;">single replacement</p>
19	$2\text{PF}_5 \rightarrow 2\text{P} + 5\text{F}_2$ <p style="text-align: center;">decomposition</p>

20	$\text{WX} + \text{Y} \rightarrow \text{WXY}$ <p style="text-align: right;">synthesis</p>
21	$\text{C}_{16}\text{H}_{32} + 24\text{O}_2 \rightarrow 16\text{CO}_2 + 16\text{H}_2\text{O}$ <p style="text-align: center;">combustion</p>

23	$\text{PKLM} \rightarrow \text{PL} + \text{K} + \text{M} \quad \text{decomposition}$
24	$\text{Be}(\text{ClO}_3)_{2(\text{AQ})} + \text{MgCrO}_{4(\text{AQ})} \rightarrow \text{Mg}(\text{ClO}_3)_{2(\text{AQ})} + \text{BeCrO}_{4(\text{S})}$ <p style="text-align: center;">Double replacement</p>

24	$\text{CH}_4 + \text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{CO}_2 \quad \text{combustion}$
25	$\text{Ba}_{(\text{S})} + 2\text{NaCl}_{(\text{AQ})} \rightarrow \text{BaCl}_{2(\text{AQ})} + 2\text{Na}_{(\text{S})}$ <p style="text-align: center;">single replacement</p>