

	Type of Reaction	Skeletons, finish them, balance whenever possible, use phase symbols	What tables did you use?
example	SR	$\text{Mg}_{(s)} + \text{H}_2\text{SO}_{2(aq)} \rightarrow \text{MgSO}_{4(aq)} + \text{H}_{2(g)}$ (balanced as is)	Periodic Table (PT), J, F
1		$\text{C}_3\text{H}_{8(g)} + \text{O}_{2(g)} \rightarrow$	
2		$(\text{NH}_4)_2\text{S}_{(aq)} + \text{Al}(\text{NO}_3)_3(aq) \rightarrow$	
3		$\text{Ba}_3\text{P}_{2(s)} \rightarrow$	
4		$\text{Al} + \text{O}_{2(g)} \rightarrow \text{Al}_2\text{O}_{3(s)}$	
5		$\text{Na}_2\text{SO}_{4(aq)} + \text{SrCl}_{2(aq)} \rightarrow$	
6		$\text{Fe}_{(s)} + \text{CuCl}_{2(aq)} \rightarrow \text{FeCl}_{3(aq)} + \text{Cu}_{(s)}$	
7		$\text{C}_8\text{H}_{18} + \text{O}_{2(g)} \rightarrow$	
8		$\text{Co}_{(s)} + \text{S}_{(s)} \rightarrow \text{Co}_2\text{S}_3(s)$	
9		$\text{KNO}_{3(s)} \rightarrow \text{KNO}_{2(s)} + \text{O}_{2(g)}$	
10		$\text{Cl}_{2(g)} + \text{LiBr}_{(aq)} \rightarrow$	
11		$\text{Al}(\text{NO}_3)_3(aq) + (\text{NH}_4)_2\text{SO}_4(aq) \rightarrow$	
12		$\text{Pb}(\text{NO}_3)_2(aq) + \text{LiF}_{(aq)} \rightarrow$	
14		$\text{BeF}_{2(s)} \rightarrow \text{Be}_{(s)} + \text{F}_{2(g)}$	

	Type of Reaction	ANSWERS	What tables did you use?
example	SR	$\text{Mg}_{(s)} + \text{H}_2\text{SO}_{2(aq)} \rightarrow \text{MgSO}_{4(aq)} + \text{H}_{2(g)}$ (balanced as is)	PT, J, F
1	combustion	$\text{C}_3\text{H}_{8(g)} + 5\text{O}_{2(g)} \rightarrow 3\text{CO}_{2(g)} + 4\text{H}_2\text{O}_{(g)}$	PT
2	double replacement	$3(\text{NH}_4)_2\text{S}_{(aq)} + 2\text{Al}(\text{NO}_3)_3(aq) \rightarrow 6\text{NH}_4\text{NO}_{3(aq)} + \text{Al}_2\text{S}_{3(s)}$	E, F
3	decomposition	$\text{Ba}_3\text{P}_{2(s)} \rightarrow 3\text{Ba}_{(s)} + 2\text{P}_{(s)}$	PT
4	synthesis	$4\text{Al} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$	PT
5	double replacement	$\text{Na}_2\text{SO}_{4(aq)} + \text{SrCl}_{2(aq)} \rightarrow 2\text{NaCl}_{(aq)} + \text{SrSO}_{4(s)}$	PT, E, F
6	single replacement	$2\text{Fe}_{(s)} + 3\text{CuCl}_{2(aq)} \rightarrow 2\text{FeCl}_{3(aq)} + 3\text{Cu}_{(s)}$	PT, F
7	combustion	$2\text{C}_8\text{H}_{18} + 25\text{O}_{2(g)} \rightarrow 16\text{CO}_{2(g)} + 18\text{H}_2\text{O}_{(g)}$	PT
8	synthesis	$2\text{Co}_{(s)} + 3\text{S}_{(s)} \rightarrow \text{Co}_2\text{S}_{3(s)}$	PT
9	decomposition	$2\text{KNO}_{3(s)} \rightarrow 2\text{KNO}_{2(s)} + 3\text{O}_{2(g)}$	PT
10	single replacement	$\text{Cl}_{2(g)} + 2\text{LiBr}_{(aq)} \rightarrow \text{Br}_{2(l)} + 2\text{LiCl}_{(aq)}$	PT, F
11	double replacement	$2\text{Al}(\text{NO}_3)_3(aq) + 3(\text{NH}_4)_2\text{SO}_4(aq) \rightarrow \text{Al}_2(\text{SO}_4)_3(s) + 6\text{NH}_4\text{NO}_3(aq)$	E, F
12	double replacement	$\text{Pb}(\text{NO}_3)_2(aq) + 2\text{LiF}_{(aq)} \rightarrow \text{PbF}_{2(s)} + 2\text{LiNO}_3(aq)$	PT, E, F
14	decomposition	$\text{BeF}_{2(s)} \rightarrow \text{Be}_{(s)} + \text{F}_{2(g)}$ (balanced as is)	PT