


14.56 mL	0.450 g	1.0000008 cm	0.0000008 cm	0.45 g	6.36×10^4 kg	300 ounces
6.5×10^{21} atoms	100 yards	100. yards	100.0 yards	4500.1 grams	2.4000 meters	2.4001 meters
The quotient of 4.56 g & 0.23 cm^3	4507 joules	5.5556×10^2 kPa	110 atm	101 atm	1.01 atm	200.59 AMU
0.457 joules	4570 joules		0.11 atm	0.110 atm	15 meters	1.5 meters
0.4570 joules	5.000×10^9 grams	1.0 cm	0.1 cm	0.10 cm	15.0 meters	1.50 meters
7°C	14°C	14.000°C	14.000004°C	7.00×10^3 °C	0.150 meters	0.15 meters
The density of a metal that has a mass of 333.55 g & 23.80 mL volume.		The % Error when your measure is $234,560 \text{ cm}^3$ and the actual value is $225,000.0 \text{ cm}^3$		The Percent Error when you measure 225 pounds but the teacher's mass is really 220. pounds		0.1500 meters
45,678,900.000 m	2.394000000 cm	25°C	0.25°C	25.0°C	25.99°C	1.0×10^{43} joules
Calculate density of an unknown metal with 76.12462 g, with volume of 14.300 mL.			What metal do you have there at left? What is the name and symbol? What is the atomic number?		Write the density formula to start. The density of pure water is 1.000 g/mL. What volume does 239 grams of water have? (do the math on back)	
Which has a greater density, 20.0 grams of aluminum, or 14.0 grams of aluminum? Are you sure?			Write the chemistry symbols for gold and for water. Label one an <u>element</u> , and the other a <u>compound</u> .		Write the density formula to start. What volume does 239 grams of gold have? Table S has the real density.	