

SF Practice

Indicate how many SF are in each measure, or do some math. name _____

14.56 mL	0.450 g	1.0000008 cm	0.000008 cm	0.45 g	6.36×10^4 kg	300 ounces
6.5×10^{23} atoms	100 meters	100. meters	100.0 meters	4500.1 grams	2.4000 meters	2.4001 meters
The quotient of 4.56 g & 0.23 cm^3	4507 joules	5.5556×10^2 kPa	110 atm	101 atm	1.01 atm	200.59 AMU
0.457 joules	4570 joules	0.1 amt	0.11 atm	0.110 atm	15 meters	1.5 meters
0.4570 joules	5.000×10^9 grams	1.0 cm	0.1 cm	0.10 cm	15.0 meters	1.50 meters
7°C	14°C	14.000°C	14.000004°C	7.00×10^3 °C	0.150 meters	0.15 meters
The density of a metal that has a mass of 333.55 g & 23.80 mL volume.		The % Error when your measure is $234,560 \text{ cm}^3$ and the actual value is $225,000.0 \text{ cm}^3$		The Percent Error when you measure 214 pounds but the teacher's mass is really 200. pounds		0.1500 meters
45,678,900.000 m	2.394000000 cm	25°C	0.25°C	25.0°C	25.99°C	1.0×10^{43} joules

<p>Calculate density of an unknown metal with mass of 76.12462 g, and volume of 14.300 mL.</p>	<p>What metal do you have there at left? What is the name and symbol? What is the atomic number?</p>	<p>Write the density formula to start. The density of pure H₂O is 1.000 g/mL. What volume does 239 grams of water have? (do the math on back)</p>
<p>Which has a greater density, 20.0 grams of aluminum, or 14.0 grams of aluminum? Are you sure?</p>	<p>Write the chemistry symbols for gold and for water. Which is an <u>element</u>, which is a <u>compound</u>?</p>	<p>Write the density formula. What volume does 239 grams of gold have? Table S can help with the density.</p>