Balancing Chemical Equations (handout #1) name: _____

Balance perfectly, name the type of reaction (Synth, Decomp, SR, DR, or Comb.) Many dashes will be "ones" and we don't write ones in chem, we are Number 1!

Page 1	Type of Reaction
$\_\_H_{2(G)} + \_\_O_{2(G)} \rightarrow \_\_H_2O_{(G)}$	
$\_\_Sr(OH)_{2(AQ)} + \_\_Li_2CrO_{4(AQ)} \rightarrow \_\_SrCrO_{4(S)} + \_\_LiOH_{(AQ)}$	
$\_\_ZnBr_{2(AQ)} + \_\_Al_{(S)} \rightarrow \_\_AlBr_{3(AQ)} + \_\_Zn_{(S)}$	
$\_\_C_{(S)} + \_\_S_{8(S)} \rightarrow \_\_CS_{2(S)}$	
$\_\_K_{(S)} + \_\_Ni(C_2H_3O_2)_{3(AQ)} \rightarrow \_\_KC_2H_3O_{2(AQ)} + \_\_Ni_{(S)}$	
$\_\_N_{2(G)} + \_\_O_{2(G)} \rightarrow \_\_N_2O_{5(G)}$	
$\_\_P_{(S)} + \_\_Cl_{2(G)} \rightarrow \_\_PCl_{5(S)}$	
$\_\_Pb(NO_3)_{2(AQ)} + \_\_LiCl_{(AQ)} \rightarrow \_\_PbCl_{2(S)} + \_\_LiNO_{3(AQ)}$	
$\_\_Al_{(S)} + \_\_S_{8(S)} \rightarrow \_\_Al_2S_{3(S)}$	
$\_\_H_2O_{(L)} \rightarrow \_\_H_{2(G)} + \_\_O_{2(G)}$	
$\_\_Mg_{(S)} + \_\_Cl_{2(G)} \rightarrow \_\_MgCl_{2(S)}$	
$\_\_C_{15}H_{32(S)} + \_\_O_{2(G)} \rightarrow \_\_CO_{2(G)} + \_\_H_2O_{(G)}$	
$\_\_C_6H_{6(G)} + \_\_O_{2(G)} \rightarrow \_\_CO_{2(G)} + \_\_H_2O_{(G)}$	
$\_\_N_{2(G)} + \_\_H_{2(G)} \rightarrow \_\_NH_{3(G)}$	

Page 2	Type of Reaction
$\_\_Li_{(S)} + \_\_TiCl_{3(AQ)} \rightarrow \_\_LiCl_{(AQ)} + \_\_Ti_{(S)}$	
$\_\_C_2H_{6(G)} + \_\_O_{2(G)} \rightarrow \_\_CO_{2(G)} + \_\_H_2O_{(G)}$	
$\_\_Rb_{(S)} + \_\_P_{(S)} \rightarrow \_\_Rb_3P_{(S)}$	
$\_\_CH_{4(G)} + \_\_O_{2(G)} \rightarrow \_\_CO_{2(G)} + \_\_H_2O_{(G)}$	
$\_\_Na_{(S)} + \_\_I_{2(S)} \rightarrow \_\_NaI_{(S)}$	
$\_\_Rb_{(S)} + \_\_S_{8(S)} \rightarrow \_\_Rb_2S_{(S)}$	
$\_\_Al(HCO_3)_{3(AQ)} + \_\_CaCrO_{4(AQ)} \rightarrow \_\_Al_2(CrO_4)_{3(S)} + \_\_Ca(HCO_3)_{2(AQ)}$	
$\_\_Li_{(S)} + \_\_SnCl_{4(AQ)} \rightarrow \_\_LiCl_{(AQ)} + \_\_Sn_{(S)}$	
$\_\_NH_{3(G)} \rightarrow \_\_N_{2(G)} + \_\_H_{2(G)}$	
$\_\_Cs_{(S)} + \_\_N_{2(G)} \rightarrow \_\_Cs_3N_{(S)}$	
$\_\_CaCO_{3(S)} \rightarrow \_\_CaO_{(S)} + \_\_CO_{2(G)}$	
$\_\_C_{10}H_{22(S)} + \_\_O_{2(G)} \rightarrow \_\_CO_{2(G)} + \_\_H_2O_{(G)}$	
$\_\_Mn_{(S)} + \_\_ZnSO_{4AQ)} \rightarrow \_\_Mn_2(SO_4)_{7(AQ)} + \_\_Zn_{(S)}$	
$\_\_C_3H_{8(G)} + \_\_O_{2(G)} \rightarrow \_\_CO_{2(G)} + \_\_H_2O_{(G)}$	