

Write the density formula →

$= \underline{\hspace{10em}}$

There are 14 different elements in the 14 beakers. The cubes have numbers scratched into them; when you choose a pair of cubes, make sure the numbers match. Randomly choose one beaker, and measure the mass of both of the metal cubes, then the volume by displacement of both cubes. Do not round the masses, measure volume to 10th mL. Do this five times, one element at a time; don't be a hoarder. No math until you measure all masses and volumes.

On separate paper, calculate the density for each. Each time write the density formula, watch significant figures. Later you will be told which elements you used.

Compare your measured density to the actual density using percent error.

Do work on white paper, answers only below. Staple your work to this handout.

Element number	Mass in grams	Volume in mL	Measured density g/cm ³	Symbol of element	Actual density g/cm ³	Your % Error for density measurement